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Chapter 15 Water and Aqueous Systems 159 SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449) This section describes the properties of water in the liquid and solid states and explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445-447) 1.

SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)

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Chemistry I - Mr. Benjamin's Classroom

Water, however, is a liquid up to 100C, has a very high surface tension (meaning it is both cohesive and adhesive), a low vapor pressure, and arranges as a low-density solid when it freezes, meaning that frozen water floats on liquid water.

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in the real world include rubbing alcohol (70% isopropyl alcohol, 30% water) and gasoline (a mixture of hydrocarbons). 2. solvent, solute 3. When an ionic solute dissolves in water, a given ion is pulled into solution by the attractive ion-dipole force exerted by several water molecules. For example, in

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(a) Boiled water tastes flat because boiled water does not contain matter like air, carbon dioxide and other minerals, So the boiled water tastes flat.
(b) Ice at zero degree centigrade gives more cooling effect than water at 0°C because, ice at 0°C absorbs 336J per gram of energy to melt to 0°C water and hence gives more cooling effect.

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